Page 4938. Empirical formula for first compound (III) should be $C_{15}H_{11}O_2N_2Cl$ instead of $C_{16}H_{11}O_2N_2Cl$.

Page 4939. Col. 2, analysis of second compound. Empirical formula for acetyl derivative of methyl ester should be $C_{18}H_{17}ClN_2O_4$ instead of $C_{18}H_{17}ClN_2P_4$.

N. O. Brace and W. B. McCormack: Synthesis and Some Novel Reactions of α,α-Dichloroperfluoroalkyl Esters.

Page 5092. Table 1, entry at middle of line 13 should read "$Cl_3CCOOCCl_2(CF_2)_{2n}Cl$ (*F*)."

Page 5095, col. 2, line 9 (from bottom) should read "tetrachloroperfluoroöctane."

Page 5095, col. 2, line 5 (from bottom) should read "$C_8F_{16}Cl_4$: Cl, 26.0."

N. O. Brace and G. J. Mantell: Hydrophobic Compounds and Polymers from Long Chain Alkanamide-Formaldehyde Condensation Reactions

Page 5176. Ref. 3 should read "U.S. Patent 2,505,649." Ref. 4 should read "U.S. Patent 2,793,142."

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George R. Evanega, Werner Bergmann, and James English, Jr.: Some Photochemistry of 1,2,3,4,5-Pentaphenylcyclohexa-1,3-diene.

Page 14. In Table I, change 728 cm.$^{-1}$ O—O emission band of *cis*-stilbene to that of ascaridole.

Page 14. Line 24, col. 1, change "4.44" to "7.44."

Page 15. Line 11, col. 1, change ref. 13–15 to ref. 13,14.

Page 15. Line 11, col. 1, add ref. 15 to structures VIII and IX; *i.e.*, line 11 should read "VIII[15], IX[15], and maleic anhydride adduct."

Page 15. Line 16, col. 2, change "21.92" to "2.92."

Page 15. Line 29, col. 2, change "XIIa" to "XIIc."

Page 15. Line 32, col. 2, change ref. 24 to ref. 24a.

Page 15. Line 35, col. 2, change "(geminal dihedral angle, 116.3)" to "(geminal dihedral angle, 116.3°)[24b]."

Page 15. Line 39, col. 2, change "(CHC dihedral angle of 109.5°[22])" to "(HCH dihedral angle of 109.5°)[24b]".

Page 15. Ref. 24, change ref. 24 to ref. 24a; add ref. 24b, H. S. Gutowsky, M. Karplus, and D. M. Grant, *J. Chem. Phys.*, **31**, 1278 (1959).

Page 16. Line 17, col. 2, change "presence" to "absence."

E. Campaigne, N. F. Chamberlain, and B. E. Edwards: Nuclear Magnetic Spectra of 1,3,5-Trithiane and 2,4,6-Substituted 1,3,5-Trithianes.

Page 135. Because of the low solubilities of β-trithioacetaldehyde and α- and β-trithiobenzaldehyde, particularly β-trithiobenzaldehyde, in the solvents used, the signal to noise ratio in some of the n.m.r. spectra was quite unfavorable. Subsequent work in these laboratories prompted us to recheck the results for α-trithiobenzaldehyde, I, and β-trithiobenzaldehyde, II, at 60 Mc. (The spectra were obtained on a Varian Associates high-resolution n.m.r. spectrometer, Model V-4300 B at Indiana University with the help of Mr. Art Clause.) Peak positions were determined relative to internal tetramethylsilane technique. Samples were run as saturated solutions at room temperature in a number of solvents. The observed peak positions are shown in Table I. The spectra of I agree with those in the original report, but the observed spectra of II are quite different from those first reported and fall in line with those expected. The previously reported spectrum of II is apparently an artifact, in which the *a* trithiane proton signal was not observed. The peak assigned to these protons may have arisen from some aromatic contaminant in the sample. The relatively acidic character of the trithiane protons is indicated by variations, with change in solvent, of their chemical shifts relative to the aromatic protons, which show no solvent shifts.

TABLE I

N.M.R. SPECTRA OF α- AND β-TRITHIOBENZALDEHYDE

Compound	Proton	A	B	C	D
α	*a*	4.43	4.65	4.63	4.75
	e	3.86	4.38	4.55	4.50
	Ar	2.65	2.65	2.65	2.80
β	*a*	4.17	4.30	4.63	4.85
	Ar	2.65	2.65	2.65	. . .

Solvents: A, deuteriodimethylsulfoxide; B, acetone; C, sulfur dioxide; D, carbon disulfide.

S. O. Winthrop, M. A. Davis, G. S. Myers, J. G. Gavin, R. Thomas, and R. Barber: New Psychotropic Agents. Derivatives of Dibenzo[*a,d*]-1,4-cycloheptadiene.

Page 233. Table I, footnote *v*, change from $[\alpha]_D = -20.8$ to $[\alpha]_D = +20.8°$.

R. E. McClure and Alexander Ross: *S*-Alkoxymethyl and *S*-Alkylmercaptomethyl Derivative(of 2-Pyridinethiol 1-Oxide.

Page 304. Recent work in these laboratories indicates that improper infrared spectra were originally obtained for 2-pyridinethiol 1-oxide. The band at 11.92 μ is actually not found with this compound. This invalidates our conclusions that the 2-pyridinethiol 1-oxide occurs primarily in the thiol form. We would now concede that the thione form suggested by Jones and Katritzky, *J. Chem. Soc.*, 2937 (1960) is most probable. Our assignments indicating *S*-alkylation for our derivatives are probably correct, however, as there is no disagreement between various workers on this type of derivatives.

William J. Considine: Triphenylcarbethoxymethylphosphonium Chloride Dihydrate.

Page 647. In col. 2, equation 3, $(C_6H_5)_3\overset{+}{P}CH_2COO\text{-}C_2H_5,\ \overset{-}{C}12H_2O$ should read $(C_6H_5)_3\overset{+}{P}CH_2COOC_2H_5,Cl\cdot 2H_2O$. $[(C_6H_5)_3\overset{+}{P}CH_2COOH,\overset{-}{Cl}]$ should be labeled "V." Delete "VI" from C_2H_5OH. $(C_6H_5)_3\overset{+}{P}CH,\overset{-}{Cl}$ should be labeled "VI". In equation 4, $(C_6H_5)_3\overset{+}{P}{=}CHCOOC_2H_5$ should be labeled "IV."

R. L. Letsinger, J. A. Gilpin, and W. J. Vullo: Peri-Substituted Naphthalene Compounds. V. 1,8-Dilithionaphthalene.

Page 673. In col. 2, paragraph 4, line 16, for "$\lambda_{\max}^{\mathrm{EtOH}}$ 231.5 mμ (log E 4.90), 289 mμ (log E 3.96)," read "$\lambda_{\max}^{\mathrm{EtOH}}$ 248 mμ (log E 4.66), 307 mμ (log E 3.86)."

Emil Koft and Francis H. Case: Substituted 1,10-Phenanthrolines, XII. Benzo and Pyrido Derivatives.

Page 867. In col. 2, line 63, for "33 g. of product," read "3.3 g. of product."

J. N. BeMiller and Roy L. Whistler: Alkaline Degradation of Amino Sugars.

Page 1163. In col. 1, line 6 from end, for sentence 1 read "Hyaluronic acid is composed of alternating 2-acetamido-2-deoxy-β-D-glucopyranosyl units linked (1 → 4) and β-D-glucopyranosyluronic acid units linked (1 → 3).[23]"

Page 1163. We wish to point out that in our publication we inadvertently omitted the work published by A. Linker, P. Hoffman, and K. Meyer, *Nature*, **180**, 810 (1957) and A. Linker, K. Meyer, and P. Hoffman, *J. Biol. Chem.*, **235**, 924 (1960) in which the authors describe the alkaline degradation (from the reducing end) of oliogosaccharides from hyaluronic acid.

Page 1164. In col. 2, line 17 from end of Experimental, for "35" read "3500."

Robert C. Lamb and P. Wayne Ayers: 1,2,6-Tribromohexane, 5-Hexenylmagnesium Bromide, and 6-Heptenoic Acid.

Pages 1441 and 1442. (a) The first word in the second sentence of the third paragraph should be "6-Bromo-1-hexene" rather than "6-Bromo-1-hexane." (b) A line of type was left out at this point, and the text should be corrected as follows:

"6-Bromo-1-hexene has been prepared in a six-step procedure from tetrahydrofurfuryl alcohol.[3] 6-Chloro-1-hexene has been obtained by the reduction of 6-chloro-1-hexyne.[4,5]"